

Chemistry 214-001, Quantitative Analysis Laboratory

Summer 2015 Syllabus

Chem 214-001, Quantitative Analysis Laboratory (1 credit hour), May 18th through June 25th, 2015

Meets on Mondays*, Tuesdays, Wednesdays, & Thursdays 8:30 am – 11:15 am in Flanner Hall 313 (FH-313)

*will meet on one Friday, May 29th, which is a University scheduled make-up day due to Memorial Day, May 25th, on which there will be no class. Attendance is expected on the make-up day, May 29th.

Prerequisite: Chem 106/102 and 112, Chem 222/224 and 226 as well as active attendance or completion of lecture Chem 212.

Instructor: Dr. Katrina Binaku

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Office Hours: Mondays & Thursdays 4 – 6 pm,
and by a scheduled appointment.

Graduate Teaching Assistant (TA): Kathryn Renyer

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Office Hours: Tuesdays 11:30 am – 12:30 pm,
and by a scheduled appointment.

Course Objectives:

- 1) To acquaint students with some of the classical and modern techniques in analytical chemistry
- 2) To teach wet chemical lab skills, efficiency and planning of experiments
- 3) To teach critical evaluation of experimental results
- 4) To become familiar with conventional data collection in commercial and academic laboratories.

Attendance Policy: It is expected students attend every scheduled laboratory [i.e lab] session/class. It is also expected students are *on time*. Additional time will not be provided to students who are absent from lab or come late to lab. Students are allowed to attend only the section in which they are enrolled. Students must have required materials and be properly dressed to perform experiments in the lab. Make-ups for exams or pre-lab quizzes will not be given unless approved by the Instructor. Students are required to initial a sign-in sheet on each day of lab, documenting and verifying their attendance. This sheet serves as a formal record. If an absence does occur, it is the absent student's responsibility to contact the Instructor promptly.

Footwear/Clothing: Closed toe, closed heel shoes are required [no sandals, flip flops, slippers, Crocs, ballet flats, boat shoes, perforated shoes, etc.] No skin on the foot can be exposed. Long pants are recommended. Shorts and skirts are not allowed. Bare skin on the lower extremities is a safety hazard: Be advised, concentrated acids/bases will be used in most of the experiments. *Lab coats & goggles are required and must be worn at all times.* Students will be sent home if proper clothing/footwear are not worn, this counts as an absence. A safety lecture will be given the 1st day of class; this lecture is required to perform lab experiments. Students will sign a lab safety sheet acknowledging their understanding and commitment to adherence of lab safety rules/policies.

Blanket statement about "technical difficulties:" It is *strongly encouraged* that all required submissions to Sakai as well as writing & printing lab reports, opening course/data/experiment files, be done on a reliable wired internet connection [not wireless], that of which the University itself provides in the Information Commons and various computer labs on the Lake Shore Campus. Under NO circumstances will excuses of "technical difficulties" be accepted as this syllabus is stating all students should use a wired internet University computer [not wireless internet] to submit work in Sakai, write & print lab reports, open course/data/experiment files. Emailing lab reports, Sakai results, or other is not allowed in place of the required means of turning in lab

reports or required submission of items in Sakai. This list is not exhaustive and do note that any activities this course may require a computer or internet connection for should be completed using University computers with wired internet connection. Use of home internet [wired or wireless], University wireless, or public wireless is at your, the student's, own risk. It is not prohibited but as Instructor has stated in this syllabus, Instructor is not responsible for ANY technical difficulties of non-University devices [cell phone, tablet, home/work/public wireless internet or computer]. Do not submit items in Sakai using a cell phone or a tablet device as these do not count as reliable internet connection tools.

Required Materials:

- One bound (NOT SPIRAL) *laboratory notebook* such as a national-brand Composition book.
- An inexpensive *calculator* having logarithm (base 10 and e), exponential, and trig functions.
- A pair of *lab goggles* [safety glasses NOT allowed] must be worn at all times in the laboratory.
- A *Lab coat* must be worn at all times in the laboratory. It offers a layer of protection against hazards. Any color is ok, but it must be long sleeve. Amazon or the Loyola bookstore sell them.
- Chem 214-001 lab manual and handouts, all handed out on 1st day of lab [always available in Sakai].
- Non-erasable pen [scientists do not write in pencil or erasable ink]. White out is not allowed.
- Use of Sakai (frequent access, submission of experimental results, etc.)

For labs #2 & 6, it may be advantageous to bring a laptop. If deemed a distraction, Instructor or TA will request that said computer be put away. Cell phones are a distraction and should not be in use during any portion of the laboratory.

Cell phones are not a calculator substitute. Cell phones are NOT allowed for use during pre-lab quizzes, the midterm exam, or the final exam and also are not allowed to be used as a calculator during lab experiments.

Laboratory Procedures:

Instructor and TA will explain the procedures and goals for each lab experiment/assignment prior to its execution. Students will be given handouts for each lab experiment/assignment beforehand. Students are expected to read the lab experiment procedures ahead of time in order to comprehend the work and complete it safely in the laboratory. Experiment handouts will also be available on Sakai for viewing at any time. A laboratory schedule, detailing projected start/end dates for each lab experiment, pre-lab quizzes, lab report due dates, and other information will be provided on the first day of class. The schedule will be posted in Sakai and FH-313. It is also at the end of this syllabus. Any aspect is subject to change.

Lab Experiment Unknown Samples (referred to as "Unknowns"):

Each student will be assigned an unknown sample whose composition is known to at least **FOUR** significant figures. **Each student will quantify a particular analyte of interest in their unknown sample and be graded on how accurately experimental determinations reflect the unknown's true composition.** Write down the unknown # in the lab notebook AND sign for it on formal sign-up sheets provided by the TA.

For each lab experiment's unknown, students will report, VIA SAKAI, their values of each individual determination (trials), the mean concentration (or percent composition), standard deviation, and ppt associated with the overall determination. *Students will be permitted to repeat each lab experiment only once (referred to*

as a 'redo'), as time permits, in order to improve technique to potentially earn a better accuracy grade. However, in a 'redo' the student must analyze a new/different unknown sample and it must be undertaken in the period established on the laboratory schedule. In order to accomplish this, **students MUST report experimentally determined results for their unknowns via SAKAI as soon as possible!** Only after Sakai submission will an accuracy grade be calculated by the Instructor. The accuracy grade evaluates a student's experimental results in comparison to the known value. When the accuracy grade is reported to the student, he/she then may decide to repeat the lab experiment or not. *Students must submit their data in Sakai and receive an accuracy grade before a 'redo' can be attempted!* Final accuracy for an experimental unknown will be determined as the better of the two reported accuracy findings if a 'redo' is completed. Lab experiments must be completed sequentially as defined in the laboratory schedule. A student CANNOT move on to the next experiment until they have determined whether to complete a 'redo' of the previous experiment. No retro-activity of a 'redo' is allowed past the redo deadline defined in the laboratory schedule. Graded accuracy will determine about 62.36% of the course grade.

Laboratory Notebook:

Notebooks must be completed in PEN. Detailed notebook requirements are listed on pages 10-11. Notebooks must be organized but not necessarily perfect and thus can contain strikeouts. White-out is not allowed.

Students must come to lab prepared in order to optimize lab efficiency. **At the start of every NEW experiment each student must have written in their notebook:**

- 1) The date and title of the experiment and 2) An introductory paragraph summarizing the purpose of an experiment & overview that may include a very brief procedure synopsis.*

*Instructor or TA will review and initial the lab notebook in class (while a pre-lab quiz is taken). A student will not be allowed to start an experiment until the notebook has these requirements completed. Notebooks are checked at the start of each lab experiment as well as during the midterm and final exams. The notebook grade determines 4.45% of the overall course grade.

Laboratory Reports:

Lab reports must be computer generated and follow the format defined on page 8-9 of this syllabus. They are to be completed individually. Plagiarizing other students' reports (current or former), book or internet sources, or lab procedures will not be tolerated. Cite outside sources when applicable. All experimental data must be included. A lab report will always contain data from the first attempt and if applicable, a second attempt (redo) if an experiment is repeated. Graded lab reports determine 13.36% of the overall course grade.

Lab report due dates are located on the laboratory schedule. Lab reports **will not** be accepted via email. Reports must be printed and handed to the TA in lab, on the due date, within the first 15 minutes of the official lab start time (8:30 am). After 8:45 am, a lab report is considered late if it is not in the possession of the TA. If a student is not present at the beginning of class on the date a lab report is due, but comes into the laboratory at any point after the first 15 minutes of the official lab start time, their lab report is *still* considered late when turned in and there are no exceptions. If a student is present on time in lab and forgets to turn in the report on the due date, it is considered late. You cannot show your TA your report on a laptop; that does NOT count as turning in your report on time as it is not printed. If a student is absent the day a lab report is due, said student must turn in the lab report at the beginning of the next lab period and will not receive penalty.

Late lab reports will receive a 15% penalty deduction each day the report is late and result in a grade of zero if not received within one week of the due date.

To assist students in improving writing skills and address any deficiencies, the first lab report (only) may be resubmitted (revised) after the first version has been graded to receive at most ½ the lost points back. Both the original graded version and revised version must be handed in. Do discuss any questions/concerns about lab reports with the Instructor or TA.

Over the course of the semester, 7 lab experiments will be completed. Each student is required to complete all 7 lab experiments and turn in experimental data for each. Writing skills are important to explain results and other important information in the “real world,” but the Instructor realizes completing lab reports is labor intensive. Therefore, students will only write lab reports for three (3) of the seven (7) lab experiments in this course.

The following list* includes the lab experiments for which a written lab report is required:

*At the discretion of the Instructor or TA, this list can be modified at any time over the course of the semester.

- 1) Lab 1: Determination of % KHP in an Unknown
- 2) Lab 3: Assay of SO₃ by Gravimetric Determination of Sulfate (Gravimetric Analysis)
- 3) Lab 6: Polyprotic Acids (Titration of a Polyprotic Acid with a Strong Base Using a pH Meter)

Laboratory Exams:

Two written exams will cover concepts pertaining to the laboratory experiments. A Midterm exam will include **Experiments 1-3** and a Final exam will include **Experiments 4-7**. Exams will cover theory, lab technique, and related calculations. Neither exam is curved. Each exam is taken once, no ‘redo’ of exams are allowed and are not offered. Grades of exams are final unless Instructor made a grading error [which must be brought to the Instructor’s attention *the day* the graded exam is returned to the student]. See lab schedule for exam dates. Make-up exams are not given. Exam grades determine 8.91% of the overall course grade.

Laboratory Quizzes (Pre-lab Quizzes):

Before the start of each new experiment a written, 15 minute pre-lab quiz will be given regarding background, procedure, and calculations to determine student preparedness for the lab experiment. **Quizzes will be given during the first 15 minutes of lab. Thus, be punctual and always get to lab on time! If one arrives late to lab, no extra time will be given to complete the pre-lab quiz.** Quiz answers must be written in pen to receive credit. If absent on the day of a pre-lab quiz, it is the student’s responsibility to schedule an appointment with the Instructor to make up the quiz BEFORE the next lab period; otherwise, the student receives a zero (0) on the missed pre-lab quiz. Pre-lab quizzes account for 3.74% of the overall course grade.

Services for Students with Disabilities (SSWD) Policy:

Necessary accommodations will be made for students with disabilities who procure a SSWD letter. Do discuss your academic needs with the Instructor as soon as possible! However, to receive any accommodations self-disclosure, proper documentation, and registration with the SSWD office at Loyola University Chicago is required. Accommodations cannot be made until the Instructor receives proper documentation. Furthermore, accommodations are not retro-active and begin only once appropriate documentation has been received by the Instructor in a timely manner. Only those accommodations that are specifically listed in the formal SSWD letter will be provided. SSWD Policies and procedures can be found here: <http://www.luc.edu/sswd/>

Journal Article Assignment:

Referencing scientific journal articles is an important aspect of research. Each student will be required to select, read, and write a report on an article published in either *Analytical Chemistry* [ACS journal] or the *Journal of Analytical Chemistry*. This assignment will aid students in recognizing principles of analytical chemistry and the importance of method development and application. Detailed information for this assignment will be handed out and posted in Sakai. The final report, typed in MS Word, and a pdf copy of the journal article must be submitted in Sakai under *Assignments*. In person, print copies of the completed report & article will not be accepted. The due date is in the laboratory schedule. This assignment determines 4.45% of the course grade.

Academic Honesty:

Both Instructor and TA encourage students to consult one another during lab experiments and outside of class. Students can converse, brainstorm, and work through questions together but copying other students' (current or previously in Chem 214) work and presenting it as one's own is unacceptable. There is a difference between sharing knowledge and cheating. If it is determined that lab reports or other materials in this course are plagiarized or have been shared between students (current or past), no credit will be given for the work in question. Cases of suspect academic dishonesty will be handled according to University policy/guidelines. Review Loyola University Chicago's policy on Academic Integrity:
http://www.luc.edu/academics/catalog/undergrad/reg_academicintegrity.shtml

Safety Points:

Unsafe actions in the laboratory will NOT be tolerated. Each day of lab is allotted ~1 safety point. Students either earn the point, or do not. All or nothing. A student will be told when a safety infraction has been witnessed by TA/Instructor and that a safety point was deducted. This tally will be documented on the daily sign-in sheet. Safety points count towards 0.94% of the overall course grade.

Safety point deductions will occur if Instructor/TA witness unsafe behavior such as: * Coming late to lab, not wearing a lab coat, borrowing lab goggles or a lab coat, eating/drinking in the lab, chewing gum, taking goggles off in FH-313 when chemicals/glassware are still on any of the 3 lab benches (even if not your chemicals or lab bench), not wearing goggles when using/cleaning glassware, chemicals, or equipment, touching face/cell phone/personal belongings with gloves on, leaving laboratory with gloves on, not cleaning chemical spills on bench top/analytical balance/fume hood, standing or kneeling on chairs, improper disposal of chemicals, etc. *List is not exhaustive; if it is determined an action is unsafe, even if it is not listed above, student will lose a safety point.

IF LABORATORY BENCHES, ANALYTICAL BALANCES, OR OTHER EQUIPMENT IN FH-313 IS LEFT DIRTY, THE ENTIRE CLASS [all students] LOSES THE DAY'S SAFETY POINT.

Lab Clean-up: Each lab period is scheduled from 8:30 am – 11:15 am, Monday through Thursday. Students must leave the laboratory at 11:15 am. Students are REQUIRED to begin cleaning their lab bench, equipment, and chemicals, no later than 11:05 am every day of the schedule laboratory course. Students are not allowed to stay past 11:15 am to do wet chemistry under any circumstances NOR can a student gain access to the laboratory room, FH-313, outside of the scheduled class day/time in LOCUS. An exception is when TA/Instructor allow students to enter lab at 8:15 am each day to sign-in and prepare for the tasks for that day.

Grading Policy:

The established grading policy is subject to change at Instructor discretion. Please note the University uses the +/- grading scale system and it will be implemented in this course. Grades are not rounded.

Grading Category	Pts	Percent
Analytical Findings (Accuracy)**	1400	62.36%
Detailed Laboratory Reports	300	13.36%
Pre-Lab Quizzes	84	3.74%
Lab Notebook	100	4.45%
Journal Article Assignment	100	4.45%
Safety points	21	0.94%
Social justice activity (Sakai, wksheet)	40	1.78%
Midterm (100), Final (100) exams	200	8.91%
Total	2245	100.0%

Grade Assignment:

Points Range	Letter Grade
2020 - 2245	A- to A
1794 - 2019	B-, B, or B+
1572 - 1793	C-, C, or C+
1347 - 1571	D-, D, or D+
Below 1347	F

**7 labs @ 200 points each

Typical Grading Scale* (%): A 100-94.0, A- 93.9-90.0, B+ 89.9-86.9, B 86.8-83.0, B- 82.9-79.9, C+ 79.8-77.0, C 76.9-72.9, C- 72.8-70.0, D+ 69.9-67.9, D 67.8-63.0, D- 62.9-60.0, F ≤ 59.9

*subject to change at the discretion of Instructor. There is no curve & final grades are not rounded.

Lab Report and Notebook Grading Rubrics:

The following is a guide of lab report/ notebook grading. Point redistribution at the discretion of the Instructor and TA is possible if deemed necessary.

Lab Report	Points
Title Page	5
Introduction/Purpose	15
Procedure	15
Results	35
Conclusion	20
Grammar/Formatting/Spelling	10
TOTAL	100

Notebook (Pts breakdown based on 7 experiments)	Points
Table of Contents	7
Title of experiment (1pt/experiment)	7
Introduction (must be signed, 2pts/experiment)	14
Results/ Raw Data and Calculations (7pts/experiment)	49
Conclusion (3pts/experiment)	21
Organization (sections labeled, writing legible)	2
Total	100

Social Justice in the Sciences:

One of the emphases of the Jesuit community is social justice. How can social justice be integrated or thought about in terms of the field of science? We will ponder this question in-class during a discussion AND in an open Forum in Sakai, with some prompts and briefing in the content. An initial inquiry will be discussed the first day of class. Two (2) forum posts total in Sakai relating to the topics at hand will be required for each student. The Forum will be open all semester and close at 11:15 am on the last day of class, June 25, 2015. Each student's post is worth 10 points, 20 points total for the Sakai forum. Towards the end of the semester [see laboratory schedule] an in-class presentation by the Instructor, followed by an in-class worksheet (20 points) students complete, will be accomplished. This work counts towards 1.78% of the overall course grade. Here are a couple of resources that you may use to engage your thinking; use this information to brainstorm potential injustices pertaining to the sciences. Additional information will be in the Sakai Forum. This activity is meant to be an enlightening, interactive experience full of open conversation on the topic. Do speak your mind in these posts but do not be disrespectful to your classmates if you disagree with their opinions.

<http://blogs.luc.edu/socialjustice/social-outreach-resources/>

<http://jesuits.org/whatwedo?PAGE=DTN-20130520124035>

IDEA (Individual Development and Educational Assessment):

IDEA is the course/instructor evaluation system that Loyola University Chicago utilizes. *Essential* and *Important* objectives have been selected by the Instructor which represent the goals and development to be achieved throughout and as a result of completing the course.

Essential objectives:

3. Learning to apply course material (improve thinking, problem solving, making decisions)
4. Developing specific skills, competencies, and points of view needed by professionals in the field most closely related to this course
11. Learning to *analyze* and *critically evaluate* ideas, arguments, and points of view

Important objectives:

9. Learning how to find and use resources for answering questions or solving problems
12. Acquiring an interest in learning more by asking questions and seeking answers

Towards the end of the semester, an email will be sent to you by an IDEA administrator, requesting the completion of the IDEA course/instructor evaluation for Chem 214-001. The objectives will be discussed the first day of lab.

See the following pages for Lab Report and Notebook Requirements!

Lab Report Format and General Guidelines: Chem 214-001

Lab reports for Quantitative Analysis are more detailed, complete than those in General or Organic Chemistry. This is an upper division lab course, and more thoroughness is expected. The lab report is a **VERY IMPORTANT** part of a laboratory based course, both at the undergraduate and graduate level.

Basic formatting: 12pt Times New Roman font, double-spaced, and out of the spirit of sustainability do feel free to print the lab reports double-sided. Define each section of the lab report in **bold (Introduction, etc.)** with the respective element names described below.

Lab reports must consist of the following elements:

Title page – lab experiment name and number centered on the page; your name, lab partners name (for partner labs only), unknown #, and date the report is due should be in the lower right corner of the page

Introduction/Purpose – brief statement of the reason for performing the experiment and the goal of the work. Then, expand on the chemistry principles. Introduce what is being learned and what will be accomplished as a result of the lab experiment. Should be no more than 2 pages

Procedure – a list of all the steps necessary to perform the lab, including any changes that may have been made to the original printed procedure.

- This can be summarized from the lab procedure but must be rewritten in one's own words! **Do not plagiarize.**
- It must be so clear that anyone not familiar with the lab would know exactly what to do.
- It should not contain the actual masses, volumes, etc. used by the student.
- Be careful writing preparation instructions for solutions. You will dissolve/dilute chemicals in a volume smaller than what the final volume will be and then dilute to the final volume mark.
For example: Dissolve 12 g KOH in 300 mL H₂O, dilute to 500.00 mL mark in a volumetric flask, parafilm, and invert to mix.

Results – list data obtained, such as volumes measured, weights, temperatures, in a table format

- Multiple trials are necessary to verify data as having good precision. All data must be shown, including repeat 'redo' lab data if applicable.
- Data must be represented in table format with appropriate column and row headings and include individually determined trials' values, averages (concentrations, percents, unknowns, etc.), standard deviation, ppt and other. When applicable include units in column headings i.e. "NaOH volume (mL)" or "mL of NaOH." Tables must be labeled with appropriate brief titles describing the contents within.
- Statistical analysis (average, standard deviation, Grubb's Test, ppt, etc.) of your data should also be included in this section.
- If applicable, include graphs in this section. Graphs must be labeled with a title & proper x and y axes labels (including units). Graphs should be constructed in Excel or a similar program.

- If graphs/figures are included, such as spectra or chromatograms, they should be accompanied with a proper label i.e. Figure 1, and brief description directly below it.
- Include calculations in this section labeled appropriately with units, chemical identity. Properly identify what is being calculated and the trial # the calculation is being completed for.
- Include general equations corresponding to each calculation i.e. general equation for **dilutions** (see example), average, standard deviation, ppt, to name a few.
Example calculation for volume of HCl for 0.100M HCl. The calculations may be written in pen neatly so they can be read and understood.
 - Show an outline of equation being used and at least one example with your values
 - ex.: $m_1v_1=m_2v_2$ $12.0M \times (v_1)=0.100M \times (1000mL)$ $v_1=8.33$ mL
 - Please utilize leading zeros before the decimal point (0.1mL and not .1mL).
- A paragraph of the results must also be present to show the student did interpret/summarize the experimental results and the data shown in tables and graphs.

Conclusion – restatement of your results, and what the results reveal

- Include a **detailed** analysis of error (at least 3 errors) in paragraph form based on the student's own data/results. An analysis of error can also be done on theoretical errors as well, even though the student may not have made these errors in the actual experiment.
 - How does the error change the outcome (concentration higher/lower than it should be, etc.)? How does the error affect the subsequent steps in the experiment?
- How can the experiment be improved and/or made simpler?
- How can the student's technique be improved?

Additional Considerations

- Order is important for scientific work – the lab report components must follow the order listed.
- Lab reports should have page numbers, located in the bottom center of each page. Staple the lab reports before handing it in.
- Feel free to print double-sided; we are a sustainable university after all!
- All parts of the report must be typed (calculations are an exception). Use 12 point font, 1.5 lines spacing, and 1 inch margins.
- Keep entire tables on a single page. If you must split up a table, remember to include column and row headings again on the next page.
- Lab reports should have good spelling, sentence structure, etc. Do not use run-on sentences, fragments, or personal pronouns (I, we, me, etc.).
- Take the time to check the work and re-read a lab report to make sure that what is written is not only clear but also makes sense.

The following has been said:

“A student could do mediocre work and write up an excellent lab report, and the work will be thought of as wonderful. A student could do wonderful work and write it up poorly, and the work will be thought of as mediocre.”

Lab Notebook Guidelines and Grading Rubric

The notebook MUST be bound (spiral notebooks are not accepted). NOTEBOOK MUST BE COMPLETED IN PEN. Leave the first 2 pages of the notebook blank. At the top of these two pages, write TABLE OF CONTENTS. Over the course of the lab, # the pages in the notebook. In the table of contents simply write the name of each experiment. Next to the name, write the page # that the experiment starts on. The table of contents does not need to be more detailed than that.

On every day of lab work, the date should be written in the notebook at the beginning of class. This will allow you to keep track of what was completed on a particular date, including solutions prepared, experimental work and calculations.

Each of the sections of the notebook should be labeled as such using roman numerals and the section headings as displayed below.

At the start of each new experiment the following is required in the notebook at the beginning of lab (i.e. completed before coming to lab):

- I. Title of experiment, date
- II. Introduction: A paragraph synopsis/overview of what the point of the experiment is, methods (titration, precipitation, etc.) or instrumentation (if applicable) utilized in the experiment. From this short paragraph, someone reading your notebook will have a basic idea of what the experiment entails. The FIRST SENTENCE of the introduction should state the purpose/what will be discovered in the particular experiment. This paragraph can be roughly $\frac{1}{2}$ a page but no more than 1 page long.

Note: The instructor or TA will initial above sections. It is the student's responsibility to get their notebook signed as required initials will count towards notebook point value. If this section is not initialed, 1pt deduction per missing signature, per experiment.

- III. Procedure (optional)
If students find it helpful to write out the entire experiment's procedure in their own words in detail, they can do so in their notebook. It is not a requirement as students will have the printed experimental procedure to reference while completing each experiment.
- IV. Results
First, the unknown number should be clearly written at the beginning of this section. This section, as described earlier in the syllabus, should contain only calculations for solutions physically prepared in class and all observations/pertinent data generated during the experiment. This includes but is not limited to color changes (initial solution color and endpoint color in a titration for example), initial/final buret readings for all experimental

trials, balance weights for solid samples, balance #, instrument settings, etc. Values written down should have units and chemical identity accompanying them i.e. 15.05mL of NaOH. All data should be written in pen. Strikeouts are acceptable as no notebook is perfect. If alterations or changes in an experimental procedure occur, this is the section to include that information as well. If experiments required generating graphs in Microsoft Excel (or other program), print out the graphs and tape/staple them in the laboratory notebook.

V. Conclusion

Brief. Restate the purpose of the experiment and what was accomplished (one or two sentences that state the unknown number and what was quantified in the unknown). If any major errors occurred in the experiment i.e. student accidentally disposed of a sample, lost product, etc state that here as well.

Example: The purpose of this experiment was to quantify the percent sodium carbonate in an unknown sample. In unknown #12, it was determined that the unknown sample contained 39.57% sodium carbonate.

*Format Check: You can request a format check after lab experiment #1 from your TA/Instructor.

Chem 214-001, Quantitative Analysis Lab, Summer 2015 Tentative Semester Schedule

Chem 214-001 Quantitative Analysis Lab Schedule* (Summer 2015)					
Assignment Dates	Week #	Class #	Date	Lab Experiments	Proposed Tasks**
	1	1	Monday, May 18, 2015	Syllabus, Safety Lecture, Equipment Check-in, Lab 1)	Safety lecture/Lab equipment/Sig Figs/Glassware review. Prep lab #1 solns
Lab 1 Prelab Quiz		2	Tuesday, May 19, 2015	Lab 1) Determination of % KHP in an Unknown	Standardize NaOH solution w/KHP; titrate unknowns
		3	Wednesday, May 20, 2015		Titrate unknowns, complete calculations, submit to Sakai. Begin REDOs
		4	Thursday, May 21, 2015		Cont lab #1 unknown titrations, Lab #1 REDOs; Prep glassware for lab #2
	2	5	Monday, May 25, 2015	Memorial Day; NO CLASSES	No Class
Lab #1 Lab Report DUE TODAY ; Lab 2 Prelab Quiz		6	Tuesday, May 26, 2015	Lab 2) Spectrophotometric Determination of Fe & Last day to FINISH lab #1 REDOS	Must complete lab #2 in one lab session; prep standards & unknown, analyze via Spec 20, then clean all glassware; last day for lab #1 REDOs
		7	Wednesday, May 27, 2015		Start lab #2 REDOs OR begin to heat/weigh empty crucibles for lab #3
Lab 3 Prelab Quiz		8	Thursday, May 28, 2015	Lab 3) Assay of SO₃ by Gravimetric Analysis of Sulfate	Lab #2 REDOs OR Prep lab #3 unknown samples & digest for 1.5hrs if time
Makeup day for Memorial holiday; attendance EXPECTED		9	Friday, May 29, 2015	Last day to FINISH lab #2 REDOS	Digest samples 1.5 hrs if not done OR filter digested samples. Put crucibles in oven w/product and leave until next lab
	3	10	Monday, June 01, 2015		Heat/weigh crucibles with product. CLEAN out crucibles; start lab # REDOs
		11	Tuesday, June 02, 2015		CLEAN out crucibles. Cont lab #3 REDOs
Lab Midterm Exam		12	Wednesday, June 03, 2015	Midterm Exam / Notebook Check #1 / Cont. lab work	Bring Calculator and Notebook! ASSAYS for labs 1-3 must be in SAKAI by MIDNIGHT TONIGHT! After midterm exam, continue lab work
Lab 4 Prelab Quiz		13	Thursday, June 04, 2015	Lab 4) Flame AA & Last day to FINISH lab #3 REDOS	Finish Lab #3 REDOs; start lab #4 digestions, prep samples for AA
Lab #3 Lab Report DUE TODAY	4	14	Monday, June 08, 2015		Cont. lab #4, interpret AA results; Prepare lab #5 iodine soln today.
Lab 5 Prelab Quiz		15	Tuesday, June 09, 2015	Lab 5) Determination of Vitamin C by Redox Titration	Must complete standardization of iodine & unknown titration in one day
		16	Wednesday, June 10, 2015	Last day to FINISH lab #4 REDOS	Complete standardization of iodine & unknown titration in one day
		17	Thursday, June 11, 2015		Cont lab #5 REDOs. Re-standardize NaOH for lab #6
Lab 6 Prelab Quiz	5	18	Monday, June 15, 2015	Lab 6) Polyprotic Acid: pH Titration Curve;	pH titration of unknown acid; graph data during lab! Cont lab #5 REDOs
Journal Article Assignment Due		19	Tuesday, June 16, 2015	Last day to FINISH lab #5 REDOS	Second pH titration of unknown acid; graph data during lab!
		20	Wednesday, June 17, 2015		Lab #6 REDOs; Prepare EDTA for lab #7
Lab 7 Prelab Quiz		21	Thursday, June 18, 2015	Lab 7) EDTA: Total Hardness as Ca, Individual [Mg] & [Ca] & Last day to FINISH lab #6 REDOs	Prep unknown dilution for IC analysis, THEN prepare EDTA solution. Standardize EDTA solution today if time allows.
Lab #6 Lab Report DUE TODAY	6	22	Monday, June 22, 2015		Standardize EDTA; Titrate unknowns & do calculations; maybe IC results
		23	Tuesday, June 23, 2015		Analyze IC results OR remake unknown dilution; continue unknown titrations; lab #7 REDOs
		24	Wednesday, June 24, 2015	Social Justice In-Class Disc. & Last day to FINISH lab #7 REDOS	Lab #7 REDOs, etc. LAST DAY OF WET CHEMISTRY.
LAB FINAL EXAM, Check-out day		25	Thursday, June 25, 2015	Final Exam / Notebook Check #2 / Equipment checkout	NO wet chemistry allowed! Bring Calculator, Notebook! ALL ASSAYS for labs 4-7 must be in SAKAI BY 11:15AM TODAY!

* This schedule is subject to change at the discretion of the instructor or TA at any point during the semester

Lab is scheduled from 8:30am to 11:15am, MTWTh. Every day attendance is expected. NO extra time will be given [cannot stay past 11:15am] nor extra days. Clean-up begins at 11:00am.

ALL LAB REPORTS ARE TO BE PRINTED OUT AND HANDED IN AT THE BEGINNING OF LAB on the DUE DATE ABOVE [within the first 15 minutes after official lab start time] ON THE DUE DATE and are late if not printed.

Emailed lab reports will NOT be accepted under any circumstances!

** Please be advised that the proposed tasks should be used as a guide and are under no circumstances the only tasks that can be performed. This is the bare minimum.

*** This schedule is meant to be a guide, to clearly map out the vigor and expectations for this course. It is not all encompassing and students must be responsible enough to keep track/stay on task.

Do note that this laboratory is designed to emphasize many important principles/concepts from the lecture course *but* the topics in lecture & lab rarely match up concurrently on a day to day basis, due to the extra detail in which lecture requires to satisfactorily cover said topics.